

¹²⁷I Pure Quadrupole Resonance Spectrum of Orthoperiodic acid, H₅IO₆

K. V. S. RAMA RAO and ALARICH WEISS

Institut für Physikalische Chemie der Universität Münster
44 Münster, Schloßplatz 4, Germany

(Z. Naturforsch. 26 a, 1813—1816 [1971]; received 12 August 1971)

The NQR-spectrum of ¹²⁷I in orthoperiodic acid, H₅IO₆, has been studied on powder samples in the temperature range 77 K ≤ T ≤ 398 K. The resonance frequencies ν_1 of the transition $m(\pm 3/2) \rightleftharpoons m(\pm 1/2)$ are 45.248 ± 0.005 MHz (77 K) and 44.976 ± 0.005 MHz (296 K). For the transition $m(\pm 5/2) \rightleftharpoons m(\pm 3/2)$ the resonance frequencies ν_2 are 83.478 ± 0.005 MHz (77 K) and 83.960 ± 0.005 MHz (296 K). The corresponding quadrupole coupling constants, $e^2 q Q/h$, are 282.977 ± 0.005 MHz (296 K), and 281.892 ± 0.005 MHz (77 K). The asymmetry parameters η have been determined as 0.2372 ± 0.0001 (296 K), and 0.2580 ± 0.0001 (77 K). A small change of 0.4% for $e^2 q Q/h$ is accompanied by a rather large change of 21% for η within the temperature range investigated. $e^2 q Q/h = f(T)$ shows a maximum at ~ 313 K. No evidence of phase transition has been observed by NQR spectroscopy between 77 K and 398 K.

Introduction

Quite strong electric field gradients act on the ¹²⁷I nuclei in covalently bonded iodides¹ and nuclear quadrupole coupling constants $e^2 q Q/h$ between 1500 and 2000 MHz are found in aliphatic and aromatic organic iodides. A factor of roughly 2 lower is the nuclear quadrupole interaction energy of pentavalent iodide, I⁵⁺, in iodates. Still lower nuclear quadrupole coupling constants have been found in periodates of the general type MeIO₄. Here the NQ-coupling constants vary between 42.410 MHz for ¹²⁷I in NaIO₄ (296 K) and 20.760 MHz in KIO₄ (296 K)²⁻⁴. Similar low coupling constants and interaction energies have been found for ¹⁸⁵Re and ¹⁸⁷Re in the perrhenates of the MeReO₄-type⁵. Comparable electric field gradients can be expected at the X-site in XO₄⁻-ions (X = I, Re), since the two ions I⁷⁺ and Re⁷⁺ have an outer shell isostructural electronic configuration.

For the periodic ion several different chemical configurations are known. Besides the tetrahedral surrounding by oxygen in IO₄⁻, an octahedral surrounding by oxygen ions (atoms) is found in the

salts of the type Me₂H₃IO₆ (Me = NH₄, Ag). The two compounds crystallize in the hexagonal system, the IO₆⁵⁻-ions having an almost regular octahedral arrangement of oxygen atoms around the iodine^{6,7}. The iodine-oxygen distance within the octahedron is 1.93 Å. Both salts undergo a phase transition^{7,8} below room temperature to antiferroelectric states. Recent studies⁹ of nuclear quadrupole interaction in (NH₄)₂H₃IO₆ by the cross-over-relaxation technique reveal that the quadrupole coupling constant, $e^2 q Q/h$, of ¹²⁷I is 25.35 MHz at room temperature and the asymmetry parameter η is zero. The discussion of the NQR results in the periodates and perrhenates of the MeXO₄-type (CaWO₄-structure)^{4,5} gave evidence that in all these isostructural compounds the relatively large quadrupole coupling probably results from contributions to the electric field gradient tensor from both the point charges $q(e)$ in the tetragonal lattice and from partial covalent bonding $q(p)$ in the ion which is distorted from tetrahedral symmetry. It has been found that both $q(e)$ and $q(p)$ are very sensitive to the changes in the oxygen parameter.

Reprints request to Prof. Dr. A. WEISS, Westfäl. Wilhelms-Universität, Institut für Physikal. Chemie, D-4400 Münster (Westf.), Schloßplatz 4.

¹ E. SCHEMP and P. J. BRAY, Physical Chemistry (An advanced treatise), Edited by HENRY EYRING, DOUGLAS HENDERSON, and WILHELM JOST, Vol. 4, 521 [1970].

² W. WEYRICH, Dissertation, Darmstadt 1971.

³ A. WEISS and W. WEYRICH, Z. Naturforsch. 24 a, 474 [1969].

⁴ M. T. ROGERS and K. V. S. RAMA RAO, Presented at the 156th National ACS Meeting, Atlantic City, N. J. 1968 (to be published).

⁵ M. T. ROGERS and K. V. S. RAMA RAO, J. Chem. Phys. 49, 1229 [1968].

⁶ L. HELMHOLTZ, J. Amer. Chem. Soc. 59, 2036 [1937]. — H. GRÄNISCHER, R. KIND, W. M. MEIER, and W. PETTER, Helv. Phys. Acta 41, 843 [1968].

⁷ H. GRÄNISCHER, W. M. MEIER, and W. PETTER, Helv. Phys. Acta 27, 216 [1954].

⁸ P. BÄRTSCH, Helv. Phys. Acta 18, 267 [1945].

⁹ R. KIND and H. GRÄNISCHER, J. Phys. Soc. Japan, Suppl. 28, 109 [1970].



Dieses Werk wurde im Jahr 2013 vom Verlag Zeitschrift für Naturforschung in Zusammenarbeit mit der Max-Planck-Gesellschaft zur Förderung der Wissenschaften e.V. digitalisiert und unter folgender Lizenz veröffentlicht: Creative Commons Namensnennung-Keine Bearbeitung 3.0 Deutschland Lizenz.

Zum 01.01.2015 ist eine Anpassung der Lizenzbedingungen (Entfall der Creative Commons Lizenzbedingung „Keine Bearbeitung“) beabsichtigt, um eine Nachnutzung auch im Rahmen zukünftiger wissenschaftlicher Nutzungsformen zu ermöglichen.

This work has been digitalized and published in 2013 by Verlag Zeitschrift für Naturforschung in cooperation with the Max Planck Society for the Advancement of Science under a Creative Commons Attribution-NoDerivs 3.0 Germany License.

On 01.01.2015 it is planned to change the License Conditions (the removal of the Creative Commons License condition "no derivative works"). This is to allow reuse in the area of future scientific usage.

It seemed worthwhile to us to study the NQR-spectrum of ^{127}I in crystalline H_5IO_6 . Here the different IO_6 -octahedra are connected to a three-dimensional frame by hydrogen bonds. Detailed crystal structure studies by X-ray and neutron diffraction^{10, 11} analysis showed that H_5IO_6 consists of slightly deformed oxygen octahedra with iodine as the central atom.

Orthoperiodic acid belongs to the monoclinic crystal class, space group $\text{P}2_1/\text{n}$ with $a = 5.28 \text{ \AA}$, $b = 18.35 \text{ \AA}$, $c = 4.95 \text{ \AA}$, $\beta = 111.4^\circ$ and $Z = 4$. The molecule H_5IO_6 consists of a slightly deformed octahedron with iodine near the centre. Five of the oxygen atoms are directly linked to hydrogens. For these five oxygen atoms the distance to the central iodine atom is 1.89 \AA . For the remaining oxygen atom, the distance $\text{I}-\text{O}$ is 1.78 \AA . It has also been suggested, that the bonds within the molecule may be covalent¹¹. This gives an indication of the relatively large quadrupole coupling [when compared to $(\text{NH}_4)_2\text{H}_3\text{IO}_6$] at the iodine site resulting from the contribution of both $q(\text{e})$ and $q(\text{p})$. This has indeed been the case, and in this paper we report the values of quadrupole coupling constants obtained at several temperatures by direct observations of ^{127}I pure quadrupole resonance spectra in polycrystalline samples of H_5IO_6 .

Experimental

The resonances reported here were observed using a frequency modulated, externally quenched superregenerative oscillator¹² and an oscilloscope. The quench frequencies were around 50 kHz. No optimum performance of the oscillator can be reached using only one tank coil for the measurement of ν_1 and ν_2 .

To measure the two transitions

$$\nu_1 = m(\pm 3/2) \rightleftharpoons m(\pm 1/2)$$

and

$$\nu_2 = m(\pm 5/2) \rightleftharpoons m(\pm 3/2)$$

at the same temperature T_a , the oscillator was first adjusted for the low frequency line (ν_1). Then a parallel coil (a shunting coil) of appropriate inductance was soldered at the top of the tuning condenser keeping the main coil in the dewar at $T_a = T$. In this way ν_1 and ν_2 could be detected at the same temperature T . The intensity of ν_2 was reduced slightly due to loss of power and filling factor. However, this did not greatly affect our measurements since the signals were fairly strong (at 296 K $S/N = 7$ for ν_1 , and $S/N = 6$ for ν_2 with the shunting coil).

The frequencies were measured by using a Rhode und Schwarz, WIK BN 4221 type frequency meter; the error in frequency measurements is $\pm 5 \text{ kHz}$ and is solely due to the line width of the NQR-lines. The temperatures between room temperature and 115 K were achieved by immersing the sample tube in a dewar containing isopentane and slowly adding liquid nitrogen with constant stirring. The melting points of petroleum ether (184 K) and carbon tetrachloride (250 K) were used as a check.

For the measurement of ν_1 and ν_2 at the boiling point of liquid nitrogen, the sample was first precooled to 115 K, using an isopentane liquid nitrogen slush bath. Then, after the sample had reached the temperature of 115 K, it was immersed in liquid N_2 . In this way a considerable improvement in signal to noise ratio was reached in comparison to samples cooled directly from room temperature to liquid N_2 .

The temperatures above room temperature were achieved by immersing the sample in paraffin and heating it with a temperature controlled oil bath. The resonances were observable to 398 K after which they began broadening since the sample started melting. Temperatures were measured with a copper constantan thermocouple and were accurate to ± 1 degree. The periodic acid (H_5IO_6 pro analysi) in crystalline form was obtained from E. Merck, Darmstadt.

A plot of the temperature variation of the ^{127}I -NQR-frequencies in H_5IO_6 is given in Fig. 1.

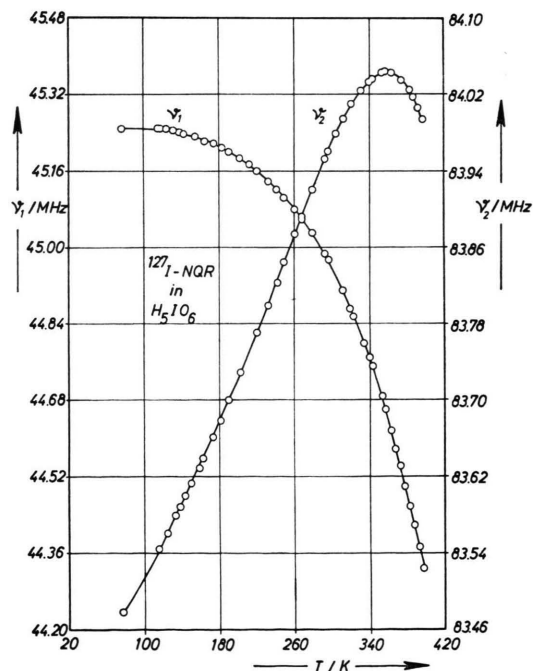


Fig. 1. NQR-frequencies $\nu_1 (m = \pm 1/2 \rightleftharpoons m = \pm 3/2)$ and $\nu_2 (m = \pm 3/2 \rightleftharpoons m = \pm 5/2)$ of ^{127}I in H_5IO_6 as a function of temperature.

¹⁰ Y. D. FEIKEMA, Acta Cryst. **14**, 315 [1961].

¹¹ Y. D. FEIKEMA, Acta Cryst. **20**, 765 [1966].

¹² H. HUTHMANN, Diplomarbeit, Münster 1971.

Results and Discussion

The quadrupole coupling constant ($e^2 q Q/h$) and the asymmetry parameter η were calculated using the tables of LIVINGSTON and ZELDES¹³. The plots are shown in Figs. 2 and 3. The values of $e^2 q Q/h$ and η calculated at a few different temperatures are given in Table I, together with the resonance frequencies ν_1 and ν_2 . Within the temperature range 77 K to 398 K, η changes by 21% whereas the $e^2 q Q/h$ varies by only 0.4%. This clearly shows that the spin 5/2 energy level scheme is more dominated by changes in η and only slightly by the maximum principal axis of the electric field gradient system. A similar effect has been observed by SEGEL and ANDERSON¹⁴ on $\text{Re}_2(\text{CO})_{10}$ where the asymmetry parameter of ^{187}Re varies by 43% whereas the coupling constant changes by only 6% in the temperature range 77 K to 340 K.

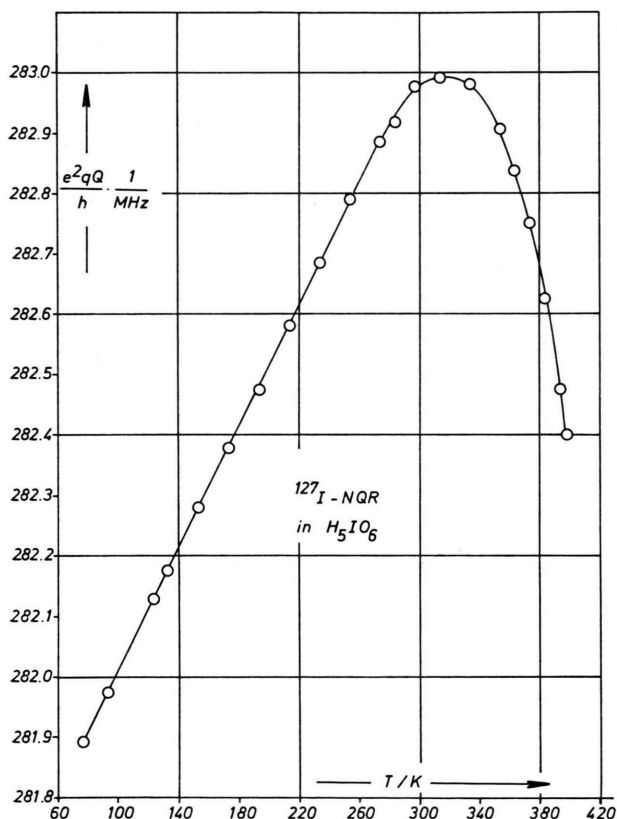


Fig. 2. The nuclear quadrupole coupling constant $e^2 q Q/h$ of ^{127}I in H_5IO_6 as a function of temperature.

Since our present work included only powder data, no information was obtained concerning the direction cosines of the electric field gradient tensor at the site of the iodine nucleus. However, from knowledge of the crystal structure, it may be assumed that q_{zz} is along the short I—O bond (1.78 Å) within the IO_6 octahedron.

In $(\text{NH}_4)_2\text{H}_3\text{IO}_6$ the IO_6 -octahedron is almost regular. Therefore $q(e)$, caused by the oxygen atoms within the first coordination sphere around the iodine, should be small and certainly smaller than in H_5IO_6 . The large difference in the coupling constants between H_5IO_6 (282.977 MHz at 296 K) and $(\text{NH}_4)_2\text{H}_3\text{IO}_6$ (25.35 MHz at 296 K) is probably caused mainly by the difference in $q(p)$. For the ammonium salt $q(p)$ should be small due to the high symmetry of the IO_6 -ion, whereas in H_5IO_6 , $q(p)$ can reach a comparatively high value.

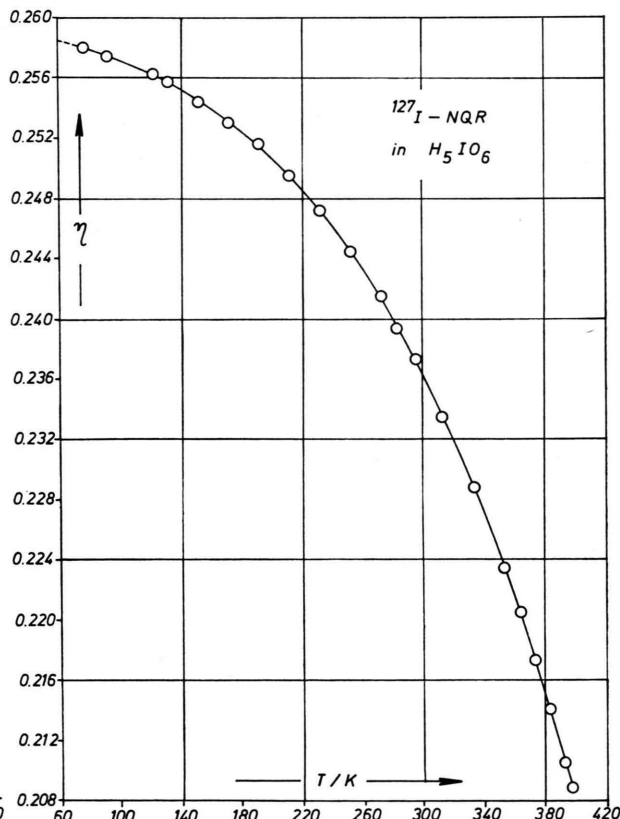


Fig. 3. The temperature dependence of asymmetry parameter $\eta = (q_{xx} - q_{yy})/q_{zz}$ at the site of the ^{127}I nucleus in H_5IO_6 .

¹³ R. L. LIVINGSTON and H. ZELDES, Tables of Eigenvalues for Pure Quadrupole Spectra, Spin 5/2, ORNL-1913 [1955].

¹⁴ S. L. SEGEL and L. A. ANDERSON, J. Chem. Phys. **49**, 1407 [1968].

Table 1.

T/K	ν_1/MHz $m(\pm 3/2)$ $\pm m(\pm 1/2)$	ν_2/MHz $m(\pm 5/2)$ $\pm m(\pm 3/2)$	η	$e^2 q Q/h$ MHz
77	45.248 ± 0.005	83.478 ± 0.005	0.2580 ± 0.0001	281.892 ± 0.005
123	45.247 ± 0.005	83.561 ± 0.005	0.2562 ± 0.0001	282.128 ± 0.005
153	45.232 ± 0.005	83.619 ± 0.005	0.2544 ± 0.0001	282.280 ± 0.005
213	45.173 ± 0.005	83.749 ± 0.005	0.2495 ± 0.0001	282.581 ± 0.005
253	45.097 ± 0.005	83.850 ± 0.005	0.2444 ± 0.0001	282.789 ± 0.005
296	44.976 ± 0.005	83.960 ± 0.005	0.2372 ± 0.0001	282.977 ± 0.005
313	44.900 ± 0.005	83.994 ± 0.005	0.2334 ± 0.0001	282.992 ± 0.005
363	44.616 ± 0.005	84.043 ± 0.005	0.2204 ± 0.0001	282.838 ± 0.005
398	44.327 ± 0.005	83.994 ± 0.005	0.2087 ± 0.0001	282.401 ± 0.005

The small influence of $q(e)$ in periodates can be understood from the fact that for a I^{7+} -ion the Sternheimer factor γ_∞ should be small compared to I^- . Also for I^- in "ionic" crystals such as $\text{NH}_4\text{TiI}_4 \cdot 2\text{H}_2\text{O}$ or Me_2PtI_6 , the main contribution of $e^2 q Q/h$ is due to $q(p)$.

The hydrogen bond plays an important role in the intramolecular interactions within crystalline H_5IO_6 . A comparison of the ^{127}I -NQR spectra in D_5IO_6 and H_5IO_6 will be quite interesting in this respect.

Nuclear Spin Relaxation by Defect Mechanisms of Self-Diffusion in Crystals

D. WOLF

Max-Planck-Institut für Metallforschung, Institut für Physik, Stuttgart

(Z. Naturforsch. **26 a**, 1816—1824 [1971]; received 30 July 1971)

The present paper shows how to incorporate the correlation between successive nuclear jumps, which are well known to play a role in defect mechanisms of self-diffusion in crystals, into the calculation of the correlation functions describing the temperature dependence of the nuclear spin relaxation in solids. The treatment is based on an extension of Torrey's theory of nuclear spin relaxation by randomly migrating nuclei. Explicit and detailed results are given for the mono-vacancy mechanisms of self-diffusion in an isotopically pure face-centred cubic crystal.

I. Introduction

BLOEMBERGEN, PURCELL and POUND (BPP)¹ showed that the time dependent nuclear magnetic spin-spin coupling is one of the most powerful mechanisms for bringing about spin-lattice relaxation in liquids and solids. In monoatomic solids this time dependence is due to atomic diffusion which may be characterized by the (temperature dependent) mean time of stay of a nucleus on a given lattice site, τ , henceforth called the "mean jump time".

Information about diffusional motions may be obtained from the measurement of the temperature dependence of the spin lattice relaxation time either in the laboratory frame, T_1 , where the splitting of the Zeeman-levels is due to the constant external magnetic field \mathbf{H}_0 , or in the rotating frame $T_{1\rho}$, where — in resonance — this splitting results from the rotating field with amplitude H_1 . If relaxation is observed in relaxation fields \mathbf{H}_0 or

\mathbf{H}_1 that are large compared with the internal local fields, the temperature variation of T_1 and $T_{1\rho}$ may be interpreted by means of perturbation theory. In this way one obtains for T_1 due to the magnetic dipolar interaction of like spins²

$$\frac{1}{T_1} = \frac{3}{2} \gamma^4 \hbar^2 I(I+1) [J^{(1)}(\omega_0) + J^{(2)}(2\omega_0)]. \quad (1.1)$$

Here $J^{(q)}(\omega)$ is the Fourier-transform of the correlation function²

$$G^{(q)}(\tau) = \sum_j \langle F_{ij}^{(q)}(t) F_{ij}^{(q)*}(t + \tau) \rangle_t. \quad (1.2)$$

The $F^{(q)}(t)$ are functions depending on the relative distance and orientation of two spins i and j . q may assume the values 0, 1, or 2².

For the spin-lattice relaxation time in the rotating frame, $T_{1\rho}$, LOOK and LOWE derived³

$$\frac{1}{T_{1\rho}} = \frac{3}{8} \gamma^4 \hbar^2 I(I+1) J_0(2\omega_1). \quad (1.3)$$

¹ N. BLOEMBERGEN, E. M. PURCELL and R. V. POUND, Phys. Rev. **73**, 679 [1948].

² A. ABRAGAM, The Principles of Nuclear Magnetism, Clarendon Press, Oxford 1961.

³ D. C. LOOK and I. J. LOWE, J. Chem. Phys. **44**, 2995 [1966].